

Class 6

Chemistry Chapter 3: Classification of Matter Answer sheet

I. Give symbols for the following elements:

1. Iron - Fe
2. Sodium - Na
3. Silver - Ag
4. Zinc - Zn
5. Chlorine - Cl

II. Give the chemical names:

1. CaO - Calcium Oxide
2. MgS - Magnesium Sulphide
3. Mn - Manganese
4. Kr - Krypton
5. FeS - Iron Sulphide

III. Give the Latin names:

1. Potassium - Kalium
2. Silver - Argentum
3. Lead - Plumbum
4. Tin - Stannum
5. Iron - Ferrum

IV. Classify the following into metals, nonmetals, metalloids and noble gases:

calcium, silver - metal

nitrogen, hydrogen - nonmetal

xenon, radon - noble gas

bismuth, arsenic - metalloid

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E. Give reasons:

1. FeS is not attracted by magnet nor it dissolves in carbon disulphide. Iron is an element attracted by a magnet. Sulphur is a yellow nonmetal soluble in carbon disulphide.
2. Oxygen is diatomic because its molecules are formed by grouping two atoms together. Ozone is triatomic because its molecules are formed by grouping three atoms together.

F. Define

1. Metal - refer to page 42
2. Nonmetals - refer to page 42
3. Atom - refer to the synopsis
4. Molecule - refer to synopsis

5. Metalloid - refer to page 42
6. Compounds - refer to synopsis
7. Element - refer to synopsis
8. Radioactive elements - refer to synopsis

G. Answer the following:

3. Refer to synopsis

6. i. Nitrogen ii. Carbon dioxide iii. Calcium oxide iv. Water

7. Examples:

i. Metals- Zinc, lead

ii. Nonmetals - hydrogen, oxygen

iii. Metalloids - antimony, arsenic

iv. Compounds - calcium oxide, iron Sulphide

v. Elements - sodium, chlorine

vi. Radioactive elements - radium, thorium

vii. Natural elements - iron, copper

viii. Diatomic gases - hydrogen, oxygen

ix. Polyatomic molecules - phosphorus, sulphur

H. Give one word:

1. Symbol

2. Noble gases

3. Polyatomic

4. Molecule

5. Malleability

6. Elements

7. Radioactive element
